LIGHT

• We measure the property of stars by looking at the light they emit

• Visible light, infrared, UV, radio are all types of Electromagnetic Radiation. They differ by having different frequencies → different colors

• EM Radiation is caused by accelerating electric charge (usually electrons since they are the lightest)

  Being accelerated

  Photon radiated (light)

  Electron emits light

  Electron absorbs light
Electromagnetic Force

- There is a force between any two bodies 1 and 2

\[ F = \frac{Cq_1q_2}{r^2} \]

with \( q_1 \) and \( q_2 \) being the charges and \( r \) being the distance between 1 and 2

- Both attractive and repulsive (positive and negative charge)
- Depends on the charges of the two bodies
- Decreases as the distance increases
- Is the same force everywhere in the Universe
- Stronger than Gravity but average charge usually equal 0

Electricity and Magnetism are different aspects of the same force.
Electromagnetic Force Examples
Rubbing a balloon on your hair or a cat builds up excess electrons on the balloon which can then stick to a wall, or even better, to a cat. 2 magnets pull towards each other or push away from each other depending on what “pole” faces each other. Both attractive and repulsive.

Gravity pulls me toward the center of the Earth but the repulsive electromagnetic force between like charges on my shoe’s surface and on the floor’s surface prevents me from going through the floor.
• Light is a bunch of photons $\rightarrow$ EM radiation or EM waves
• wavelength ($\lambda$) = distance between waves

• Period = time between wave peaks
• Frequency (ν) = 1/period = how rapidly wave is changing

So 60 Hz = 60 Hertz = 60 beats per second is the same as a period of 0.016 seconds

• Can use either wavelength or frequency to characterize light $\rightarrow$ frequency proportional to energy

• $E = hf = \text{constant} \times \text{frequency}$
$\rightarrow$ high frequency photons are also high energy photons
velocity = wavelength X frequency

velocity sound = 1 mile/5 seconds
velocity light = 1 mile/5 microseconds = 300,000 km/sec

Lightning occurs. You see the flash almost immediately but hear the thunder later. If you count slowly, then if thunder happens at 5 seconds the lightning strike occurred about 1 miles away.

If you see and hear lightning at the same time it has struck very close to you (my homes have been hit 3 times. In 1967, lightning hit the window well of my basement bedroom 8 ft from me at 6:00 AM and set the wall on fire. I put out the fire, went upstairs to tell my Dad who said “go back to sleep”. Later, after seeing the destroyed wall he said “I guess you were right”)

high energy
high frequency
Small wavelength
different types of electromagnetic radiation

low energy
low frequency
Large wavelength

Visible Light

Wavelength (nanometers)

Wavelength (meters)
THE ELECTROMAGNETIC SPECTRUM

Wavelength (in meters)

Common name of wave

Sources

Frequency (waves per second)

Energy of one photon (electron volts)
We observe light as continuous or discrete spectrum.

- **Continuous spectrum**: Light emitted by a blackbody is continuous across all wavelengths.

- **Absorption line spectrum**: Atoms in a gas cloud absorb specific wavelengths of light, producing a spectrum with dark lines.

- **Emission line spectrum**: Atoms in a gas cloud reemit the same wavelengths at which they absorbed light energy, producing a spectrum with bright lines.
Continuous Spectrum

- Radiation of light due only to Temperature of object
- All frequencies

- Peak of frequency spectrum depends on Temperature
  \[ \text{wavelength}_{\text{max}} = \frac{3,000,000}{T} \]
  with wavelength in nanometers and T in Kelvin

- Total energy emitted
  \[ E = \sigma \times T^4 \]
  \( \sigma = \text{constant} \)
Temperature

Temperature ↔ Velocity ↔ Energy

At higher Temps

higher velocities
more acceleration of electrons
more light emitted

• Kelvin Scale

Absolute 0 = 0^0 K
= -273^0 C
= -459^0 F

at high T Kelvin and Centigrade about same
EXAMPLES

normal, incandescent light bulbs $T=5000K$

people $T=300K$ → infrared

campfires, stoves $T=600-1000K$, start to glow red

→ Measure spectrum at a distance measures temperature of source (like a star)

high energy photons

low energy photons
Discrete Spectrum

- “spikes” at specific frequencies
- Depends on which atoms are present
- Examples include fluorescent or Neon or Mercury lights
- Can be used to identify chemical composition of objects (spectroscopy)
Atoms and Energy Levels

emission lines can tell one atom from another – in this case Hydrogen from Mercury from Neon
Atoms and Energy Levels

• An atom is a nucleus surrounded by electrons held together by the electromagnetic force

• Electron can be in different energy states; Changes in electron energy states (Quantum Leaps) produce discrete spectrum

• Transitions between different atomic energy states either emit or absorb light; the energy of the light (the photon’s frequency) is equal to the difference between the atomic energy states

• Pattern of photon frequencies tells what atom is emitting the light

\[ E(\text{photon}) = hf \]

\[ h = \text{Planck’s constant} \]

\[ f = \text{frequency} \]
Hydrogen

- Simplest atom – just one electron and one proton
- “heavy” hydrogen or deuterium adds one neutron to the nucleus
Atoms and Energy Levels - H

Absorbs light

emits light

Incoming photon, $\lambda = 656.3\ nm$

Emitted photon, $\lambda = 656.3\ nm$
Hydrogen lines

For Hydrogen, the lines in the visible spectrum are transitions to the $n=2$ (Balmer). Those to the $n=1$ are in the UV (Lyman series).
EXTRA SLIDES
How fluorescent light works

• Tube filled with Mercury and Argon gas
• Initial HV heats up gas → Argon “plasma”
• moves electrons in Mercury to higher energy levels
• electrons “fall” to lower energy levels and emit UV light
• UV light absorbed by phosphor coating on inside walls of tube and is reemitted at lower energy, with mix of colors that appears white
Gravity vs Electric Force

- electric force dominates daily life
  - all senses
  - all chemistry

- easy to observe much stronger then gravity
  - floor prevents us from falling to Earth’s center
  - can stick a balloon to the wall
  - levitating magnets

- why gravity is so weak is one of the unanswered questions in physics. Extra Dimensions? more later in course
Gravity vs Electric Force

- electric and gravitational forces \( \rightarrow \) same form
- compare strengths for electron and proton in Hydrogen

\[
\frac{F_{\text{gravity}}}{F_{\text{electric}}} (\text{Hydrogen}) = 10^{-39}
\]

\[
= 0.0000000000000000000000000000000000000000001
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